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Chapter 10 Chemical Quantities Assessment

Chapter 10 Chemical Quantities
Assessment Section 10.1 – The Mole: A
Measurement of Matter You often
measure the amount of something by
count, by mass, or by volume. A mole
(mol) of a substance is 6.02 x
1023representative particles of that
substance. 6.02 x 1023 is called
Avogadro's number.

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Chapter 10 - Chemical Quantities - 10 Assessment - Page 342: 121 Answer An atom is the smallest particle of an element that retains its identity in a chemical reaction, whereas a molecule is a group of atoms joined together by covalent bonds.

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Section 10.3 - Percent Composition and Chemical Formulas. The percent by mass (percent composition) of an element in a compound is the number of grams of the element divided by the mass in grams of the compound multiplied by 100%. % mass of element = mass of element x 100. mass of compound.

Chapter 10 - Chemical Quantities

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Section Review Objectives • Relate
Avogadro's number to a mole of a
substance • Calculate the mass of a
mole of any substance • Describe
methods of measuring the amount of

something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles.

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Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how

to calculate the mass of a mole of any substance.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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Calculations and Chemical Equations. 10.1 Equation Stoichiometry 369 The ratio of moles of P 4O 10 to moles of P (which came from the subscripts in the chemical formula, P 4O 10) provided the key conversion factor that allowed us to convert

Chapter 10 ChemiCal alCulations and equations

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